

Specific Heat

- heat added or subtracted per unit change in Temp
- constant only specific to the process

$$c = \frac{dq}{dT}$$

$$C_v = \left(\frac{dq}{dT}\right)_{\text{constant volume}}$$

$$C_p = \left(\frac{dq}{dT}\right)_{\text{constant pressure}}$$

$$\delta q = C_v dT$$

$$\delta q = C_p dT$$

constant Volume Process

$$\delta q = de + p dv \rightarrow 0$$

$$de = \delta q = C_v dT$$

$$\int_0^e de = \int_0^T C_v dT \quad \begin{array}{l} \text{Assuming} \\ \text{at} \\ T=0 \\ e=0 \\ \downarrow \\ C_v = \text{const} \end{array}$$

$$e = C_v T$$

For constant Pressure

$$\delta q = dh - v dp \rightarrow 0$$

$$\delta q = dh = C_p dT$$

$$h = C_p T$$

Assuming
@ $T=0$
 $h=0$
 $C_p = \text{const.}$

These hold for any perfect gas, and any process.

$$e = C_v T$$

$$h = C_p T$$

pg 51
ex 4.6

For air @ standard
sea-level

$$C_v = 720 \frac{\text{J}}{\text{kg}\cdot\text{K}}$$
$$= 429 \frac{\text{ft}\cdot\text{lb}}{\text{slug}\cdot\text{R}}$$

$$@ T = 288.16\text{K}$$
$$= 519\text{R}$$

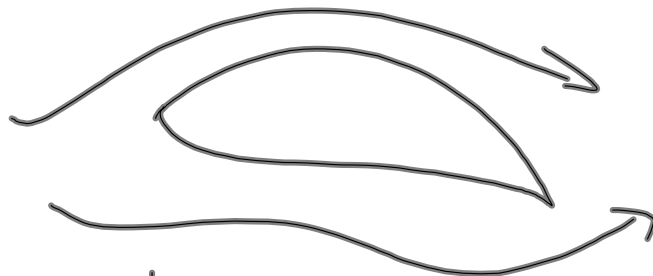
$$C_p = 1008 \frac{\text{J}}{\text{kg}\cdot\text{K}}$$
$$= 6006 \frac{\text{ft}\cdot\text{lb}}{\text{slug}\cdot\text{R}}$$

$$e = C_v T$$

$$h = C_p T$$

1. Constant Volume
2. Constant Pressure.
3. Adiabatic $\Rightarrow \delta q = 0$
 \rightarrow No heat added or taken out
4. Reversible - No friction or other external (dissipative) effects.
5. Isentropic - BOTH
 Adiabatic + Reversible

- Look at flow over an air foil



\rightarrow No external sources of heat or heat exchange.

*Note: No heat added or taken out does NOT mean No Temp Change! *

iP. Change in volume
 - compressible
 \rightarrow work

iP. Constant volume
 - incompressible
 No work, No Temp. change.

Assuming an Adiabatic Process

$$\delta q = 0$$

$$de = \cancel{\delta q} - p dV = C_v dT$$

$$dh = \cancel{\delta q} + v dp = C_p dT$$

$$\frac{de}{dh} = \frac{-p dV}{v dp} = \frac{C_v}{C_p}$$

$$\text{or } \frac{dV}{V} = -\frac{dp}{p} \frac{C_v}{C_p}$$

$$\frac{dp}{p} = -\frac{dV}{V} \frac{C_p}{C_v}$$

$$\frac{C_p}{C_v} = \gamma = \text{ratio of specific heat}$$

$$1.4 \text{ for air } \int_{v_1}^{v_2} \frac{dp}{p} = -\gamma \int_{v_1}^{v_2} \frac{dV}{V}$$

Look at pts in an isentropic flow

$$\ln \frac{p_2}{p_1} = -\gamma \ln \frac{v_2}{v_1}$$

$$\ln p_2 - \ln p_1 = -\gamma (\ln v_2 - \ln v_1)$$

$$\ln \frac{p_2}{p_1} = -\gamma \ln \frac{v_2}{v_1}$$

$$\frac{p_2}{p_1} = \left(\frac{v_2}{v_1} \right)^{-\gamma} \quad \left(\frac{2}{1} \right)^{-1} = \frac{1}{2}$$

$$\frac{p_2}{p_1} = \left(\frac{v_2}{v_1} \right)^{-\gamma} = \left(\frac{v_1}{v_2} \right)^{\gamma}$$

$$\frac{P_2}{P_1} = \left(\frac{V_1}{V_2} \right)^\gamma \quad \text{if } \nu = \text{Sp. Vol}$$

$$p = \rho R T$$

$$p v = R T$$

$$v_1 = \frac{R T_1}{P_1}$$

$$v_2 = \frac{R T_2}{P_2}$$

$$\frac{P_2}{P_1} = \left(\frac{\frac{R T_1}{P_1}}{\frac{R T_2}{P_2}} \right)^\gamma$$

$(6x)^2 = 6^2 \cdot x^2$

$$\frac{P_2}{P_1} = \left(\frac{T_1 P_2}{T_2 P_1} \right)^\gamma = \left(\frac{T_1}{T_2} \right)^\gamma \left(\frac{P_2}{P_1} \right)^\gamma$$

$$\left(\frac{P_2}{P_1} \right)^{1-\gamma} = \left(\frac{T_1}{T_2} \right)^\gamma$$

$$\frac{P_2}{P_1} = \left(\frac{T_1}{T_2} \right)^{\frac{\gamma}{1-\gamma}}$$

$$(6^2)^{1/3} = (X^3)^{1/3}$$

$$X = 6^{2/3}$$

$$\frac{P_2}{P_1} = \left(\frac{T_2}{T_1} \right)^{-\frac{\gamma}{1-\gamma}} = \left(\frac{T_2}{T_1} \right)^{\frac{\gamma}{\delta-1}}$$

$$\left(\frac{P_2}{P_1} \right) = \left(\frac{V_1}{V_2} \right)^\gamma = \left(\frac{T_2}{T_1} \right)^{\frac{\gamma}{\delta-1}} = \left(\frac{P_2}{P_1} \right)^\gamma$$

Isentropic (compressible)
Flow

Incompressible

Compressible

Bernoulli

Euler

Continuity = $A_1 \vec{V}_1 = A_2 \vec{V}_2$
(V = Velocity)

$$\rho_1 A_1 \vec{V}_1 = \rho_2 A_2 \vec{V}_2$$

Isentropic
Flow eq.

Ex 4.7 Sea level!

$$T_1 = 250 \text{ K}$$

$$p_1 = ?$$

$$p_\infty = 1.01325 \cdot 10^5 \frac{\text{N}}{\text{m}^2}$$

$$T_\infty = 288.16 \text{ K}$$

$$\frac{p_1}{p_\infty} = \left(\frac{T_1}{T_\infty} \right)^{\frac{\gamma}{\gamma-1}} = 1.4$$

$$p_1 = 61,400 \frac{\text{N}}{\text{m}^2}$$